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Key Concepts: Terms in this set (25)
The coefficients in a chemical equation represent the
(a) masses in grams of all reactants and products.
(b) relative number of moles of reactants and products.

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To understand stoichiometry, start with this introduction to the topic. It might also help to review

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molecules and moles, which includes how chemical formulas work. Ready for another quiz? Here's a quick self-test about the mole. If you'd rather switch gears, see if you know the answers about how chemistry explains the real world.

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d. 12% e. 2% 11. How many grams of H₂O will be
formed when 32.0 g H₂ is allowed to react with 16.0
g O₂ according to $2 \text{H}_2 + \text{O}_2 \rightarrow 2 \text{H}_2\text{O}$ a. 9.00 g b.

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16.0 g c. 18.0 g d. 32.0 g e. 36.0 g 12. When 2.00 g of H_2 reacts with 32.0 g of O_2 in an explosion, the final gas mixture will contain: a. H_2 , H_2O , and O_2 b. H_2 and H_2O only c. O_2 ...

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