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~~24 ChemQuest 29 Monc LcW s Stnuctunes tr I Name Date: Hour: Introduction Questions Draw the Lewis structures for the following molecules. Part c is done for you. a) H<sub>2</sub>O (oxygen is central) b) SO<sub>2</sub> c) Nz~~

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It t-t (a [Jrt @ N=N @ @ f\ @ 2 In question 1 above you should see three different kinds of bonds of bonds as best you can. a) single bond: one electron (e<sup>-</sup>) share b) \*-1; double bond:

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24 ChemQuest 8 Name: \_\_\_\_\_ Date: \_\_\_\_\_ Hour: \_\_\_\_\_ Information: Structure of the Atom Note the following symbols: (they are not to scale) = proton (positive charge) = electron (negative charge) = neutron (no charge) The following three diagrams are hydrogen atoms: 1 1H 2 1H 3 1H The following three diagrams are carbon atoms: Notice the type of notation used for atoms: 12 13 14 6C 6 C 6C (6 ...

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included in the Lewis Structure.  $4 + 3(6) + 2 = 24$  (carbon has four and each oxygen has six; add two for the -2 charge)  $5 + 3(1) = 8$  (nitrogen has five; each hydrogen has one) Step #2: Draw the "skeleton structure" with the central atoms and the other atoms, each connected with a single bond. Step #3: Add six more electron dots to each atom except the

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Steps for Writing Lewis Structures. Example  $\backslash(\backslash\text{PageIndex}\{2\}\backslash)$  1. Determine the total number of valence electrons in the molecule or ion. Each hydrogen atom (group 1) has one valence electron, carbon (group

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14) has 4 valence electrons, and oxygen (group 16) has 6 valence electrons, for a total of  $[(2)(1) + 4 + 6] = 12$  valence electrons. 2.

10.5: Writing Lewis Structures for Covalent Compounds ...

included in the Lewis Structure.  $4 + 3(6) + 2 = 24$  (carbon has four and each oxygen has six; add two for the -2 charge)  $5 + 3(1) = 8$  (nitrogen has five; each hydrogen has one) Step #2: Draw the "skeleton structure" with the central atoms and the other atoms, each connected with a single bond.

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