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the subshell before

pairing electrons. 2.

packets/quanta of

electromagnetic

energy 3. the SI unit

of frequency 4. An

atomic orbital can

hold no more than two

electrons. 5. the

amount of energy

required to move an

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electron from its present energy level to the next higher one

6. the modern description of the location and energy of electrons in an atom

7.

~~ELECTRONS IN ATOMS~~

CHAPTER. 5

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Section. Electrons in
Atoms. What You'll
Learn. You will
compare the wave
and particle models of
light. The key to
understanding the
chemical behavior of
fire-works, and all
matter, lies in
understanding how
electrons are
arranged in Analysis
Answers will vary.

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Answer Key~~

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in this set (82)

Rutherford's model-
also called nuclear
model-did not explain
how the electrons

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were arranged

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study tools. Explain how the gaseous neon atoms in a neon sign emit light? The light of a neon sign is produced when electricity passes through a tube filled with neon gas....

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Answers~~

1) Electrons circulate

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Around the nucleus 2)

Electrons can jump from one level to the next but can NOT

exist between levels

(ladder) 3) Energy

levels are NOT evenly spaced (escalator) 4)

Energy lost or gained is NOT always the

same 5) The higher the energy level, the

easier for the

electrons to leave

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Vocabulary

~~Honors Chemistry~~

~~Chapter 5 – Electrons~~

~~in Atoms (Review ...~~

the formula $2n^2$. the
maximum number of
electrons that can
occupy an energy
level. opposite spin. in
order to occupy the
same orbital 2
electron must have.
filled energy sub

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levels. stable electron configurations are likely to contain. all the orbitals contain one electron, with spins parallel.

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states that single electrons with the same spin must occupy each equal-

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energy orbital before additional electrons with opposite spins can occupy the same orbitals. Pauli

exclusion principle. states that a maximum of two electrons may occupy a single atomic orbital, but only if the electrons have opposite spins. photoelectric effect.

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electrons in atoms
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