

## Pearson Chapter 8 Covalent Bonding Answers

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242 Chapter 8 • Covalent Bonding Single Covalent Bonds When only one pair of electrons is shared, such as in a hydrogen molecule, it is a single covalent bond. The shared electron pair is often referred to as the bonding pair. For a hydrogen molecule, shown in Figure 8.4, each covalently bonded atom equally attracts the pair of shared electrons.

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Michelle\_McGee74 TEACHER. Pearson Chemistry Chapter 8: Covalent Bonding. Covalent Bond. Molecule. Diatomic molecule. Molecular compound. A bond formed by the sharing of electrons between atoms. A neutral group of atoms joined together by covalent bonds. A molecule consisting of two atoms.

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Covalent Bonding 8.2 The Nature of Covalent Bonding > 6 Copyright © Pearson Education, Inc., or its affiliates. All Rights Reserved. • Combinations of atoms of the nonmetals and metalloids in Groups 4A, 5A, 6A, and 7A of the periodic table are likely to form covalent bonds. • The combined atoms usually acquire a total of eight electrons, or an

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## Chapter 8

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Polar Covalent Bonds •Though atoms often form compounds by sharing electrons, the electrons are not always shared equally. •Fluorine pulls harder on the electrons it shares with hydrogen than hydrogen does. •Therefore, the fluorine end of the molecule has more electron density than the hydrogen end.

## Chapter 8 Concepts of Chemical Bonding - Central Lyon

Chapter 8 Covalent Bonding 183 Section Review Objectives • State a rule that usually tells how many electrons are shared to form a covalent bond • Describe how electron dot formulas are used • Predict when two atoms are likely to be joined by a double or a triple covalent bond • Distinguish between a single covalent bond and other covalent bonds

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