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Pt 1

Part 8 Covalent
Bonding through
drawing water
example in table

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~~Ionic Bonding~~

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Page 6/41

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~~Bonding~~ Chapter
8 *Basic Concepts
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of Chemical

Bonding: Part 3

of 8 Chapter 8

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1 of 3 ~~Lewis~~

~~Diagrams Made~~

~~Easy: How to~~

~~Draw Lewis Dot~~

~~Structures~~

Valence Bond

Theory, Hybrid

Orbitals, and

Molecular

Orbital Theory

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Made Easy The

Periodic Table:

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Energy, and Elec

tronegativity

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Allotropes of

Carbon **VSEPR**

Theory:

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240 Chapter 8 •
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The Covalent Bond
d-!) .)DEAAtoms
gain stability
when they share
electrons and
form covalent
bonds. Real-

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World Reading
Link Have you
ever run in a
three-legged
race? Each
person in the
race shares one
of their legs
with a teammate
to form a single
three-legged
team.

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and Molecular

Structure 8-3

There are two

types of

repulsive forces

between the two

atoms. First,

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the nuclei repel because they are both positively charged. Second, the electrons repel because they are both negatively charged. The attractive forces between the two atoms result from the

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**Covalent Bonding
and Molecular
Structure**

Section 8.4 -
Polar Bonds and
Molecules.

Covalent bonds
involve sharing
electrons
between atoms.
When the atoms
in the bond pull
equally, the

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Covalent

Bonding
electrons are shared equally,
and the bond is nonpolar. When
the atoms in the bond pull
unequally, the bonding
electrons are pulled closer to
one atom, and the bond is
polar.

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Answers
Covalent bond A

bond formed by

the sharing of

electrons

between atoms.

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7. When atoms
form a polar
covalent bond,
the atom with

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the greater attraction for electrons has a partial charge.

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and molecular
structure.

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Covalent Bond
Single Covalent

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Bonds (cont.) •

Atoms in group 16 can share two electrons and form two covalent bonds.

• Water is formed from one oxygen with two hydrogen atoms covalently bonded to it.

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and Change

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2 molecule more

Bonding

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Molecular

Compounds ...

It consists of

two main

sections:

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answers Review

Objectives •

State a rule that usually tells how many electrons are shared to form a covalent bond •

Describe how electron dot formulas are used • Predict

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