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Heat and temperature are different physical quantities. When two objects with different temperatures are in contact with each other, heat flows from the hotter system to the colder one. Heat is a measure of energy. An increase or decrease in mechanical energy in a system always accompanies an equal decrease or increase of heat, respectively.

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- Heat is a measure of the total kinetic energy of the particles in a body
- Heat is measured in Joules (J) or kilojoules (kJ).
- Heat energy flows from bodies of high temperature to bodies of lower temperature.
- The heat content of a body depends on its mass, temperature and the material it is made of.

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Heat, energy that is transferred from one body to another as the result of a difference in temperature. If two bodies at different temperatures are brought together, energy is transferred—i.e., heat flows—from the hotter body to the colder. The effect of this transfer of energy usually, but not always,

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is an increase in the temperature of the colder body and a decrease in the temperature of the hotter body.

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