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Problem : Given that
the free energy of
formation of liquid
water is -237 kJ / mol ,
calculate the potential
for the formation of

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hydrogen and oxygen from water. To solve this problem we must first calculate ΔG for the reaction, which is $-2 \times (-237 \text{ kJ / mol}) = 474 \text{ kJ / mol}$. Knowing that $\Delta G = -nFE^\circ$ and $n = 4$, we calculate the potential is -1.23 V .

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of thermodynamics. 1.

Based on graph P-V

below, what is the

ratio of the work done

by the gas in the

process I, to the work

done by the gas in the

process II? Known :

Process 1 : Pressure

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(P) = 20 N/m². Initial
volume (V₁) = 10 liter
= 10 dm³ = 10 × 10⁻³
m³

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of pure substances.

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chapter 05:
irreversibility and
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The first law of

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problems and

solutions. 1. 3000 J of

heat is added to a

system and 2500 J of

work is done by the

system. What is the

change in internal

energy of the system?

Known : Heat (Q) =

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+3000 Joule. Work

(W) = +2500 Joule .

Wanted: the change
in internal energy of

the system. Solution :

The equation of the
first law of
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The first law of
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Processes (Ideal Gas)

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A steady flow
compressor handles
113.3 m³ /min of
nitrogen ($M = 28$; $k =$
1.399) measured at
intake where $P_1 = 97$
KPa and $T_1 = 27$ C.
Discharge is at 311
KPa. The changes in
KE and PE are
negligible. For each of
the following

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In many courses, the instructor posts copies of pages from the solution manual. Often the solution manual does little more than show the quickest way to obtain the answer and says nothing about WHY each step is taken or

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HOW the author knew
which step to take
next.

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Chapter-10

Examples. A Carnot
vapor refrigeration
cycle is used to
maintain a cold region

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at 0 °F where the ambient temperature is 75 °F. Refrigerant R-134a enters the condenser as saturated vapor at 100 lbf/in² and leaves as saturated liquid at the same pressure. The evaporator pressure is 20 lbf/in². The mass flow rate of refrigerant is 12 lbm/s.

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Problem # 1 Since the
containers are
insulated, no heat
transfer occurs
between the gas and
the external

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Problems, and

since the gas
expands freely into
container B there is

no resistance

"pushing" against it,
which means no work
is done on the gas as
it expands.

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Second Law. 1. Two

kg of air at 500kPa,

80 °C expands

adiabatically in a

closed system until its

volume is doubled

and its temperature

becomes equal to that

of the surroundings

which is at 100kPa

and 5 °C.

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Second Law

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common

thermodynamic
equations and sample
problems showing a
situation in which
each might be used.

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PRACTICE 185

PROBLEMS FOR

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Properties 1. If an object has a weight of 10 lbf on the moon, what would the same object weigh on

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Jupiter? Jupiter...

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Properties

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hermodynamics:-Prob

lem 1:-A container

holds a mixture of

three nonreacting

gases: n_1 moles of

the first gas with

molar specific heat at

constant volume C_{v1} ,

and so on. Find the

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molar specific heat at constant volume of the mixture, in terms of the molar specific heats and quantities of the three separate gases. Concept:-

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how much you know

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Problem 1.1* The plot

of $V = V(P, T)$ for a

gas is shown in Fig.

1.1. Determine. the

expressions of the

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two second derivatives of the volume of this plot. (note: the principle curvatures of the surface are proportional to these second derivatives).

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